UNIT 7 TEST REVIEW

The Unit 7 test is on ***Friday, December 9, 2011***. In order to prepare for this test, you should review all of your notes, warm up questions, worksheets, activities we have done and quizzes along with completing all of this in detail. **Remember**, you need to actively study, **NOT** just read over this sheet multiple times. Actively studying includes doing some of these items or more: Practice problems or questions we have done, re-write notes, do the quizzes from the book online (google Glencoe Chemistry)

I will be available for questions by email – Jaime.degarmo@cfisd.net until 930 Thursday night or if you come in after 6:50 am on Friday morning. The answers will be posted on my website Thursday afternoon.

1. Label the family names, group and period numbers on the table below.

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1. Explain what is unique about the Noble Gases.
2. List the three things elements in the same group have in common.
	1. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	2. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	3. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
3. Explain Electronegativity.
4. What does the Ionization energy value tell you?
5. Label the metals, nonmetals, metalloids using colored pencils. 
6. Write the Oxidation numbers for the elements represented in the picture below.



1. Explain how Mendeleev arranged the Periodic Table.
2. Explain how Moseley re-arranged Mendeleev’s Periodic Table.
3. The state of matter of elements on the Periodic Table is at \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ temperature.
4. What are synthetic elements? And where are they located?
5. List the properties of metals and give two examples.
6. List the properties of nonmetals and give two examples.
7. List the properties of metalloids and give two examples.
8. List the four factors the trends of the periodic table depend on.
	1. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	2. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	3. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	4. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
9. Illustrate what happens to the diameter of a Potassium atom when it becomes a positively charged Potassium ion.
10. Illustrate what happens to the diameter of a Chlorine atom when it becomes a negatively charged Chlorine ion.
11. On the back of this page is a blank periodic table. Show, with arrows, the trends of the Periodic table. Be sure to label if it is increasing or decreasing in the direction of the arrow.