

Name \_\_\_\_\_

Solubility Rules\_K

1. Classify each of the substances as being soluble or-insoluble in water.

- a. potassium bromide  $\frac{KBr}{S}$
- b. lead (II) carbonate  $\frac{PbCO_3}{I}$
- c. zinc hydroxide  $\frac{Zn(OH)_2}{I}$
- d. sodium acetate  $\frac{NaC_2H_3O_2}{S}$
- e. cadmium (II) sulfide  $\frac{CdS}{I}$

- f.  $Mg_3(PO_4)_2$   $\frac{I}{S}$
- g.  $KOH$   $\frac{S}{S}$
- h.  $NiCl_2$   $\frac{I}{I}$
- i.  $Hg_2SO_4$   $\frac{I}{I}$
- j.  $PbI_2$   $\frac{I}{I}$

2. The following reactions take place in water. Fill in each parenthesis with aqueous (aq) or solid (s).

