

NO #10? #2

Name: KEY  
Period: \_\_\_\_\_ Date: \_\_\_\_\_

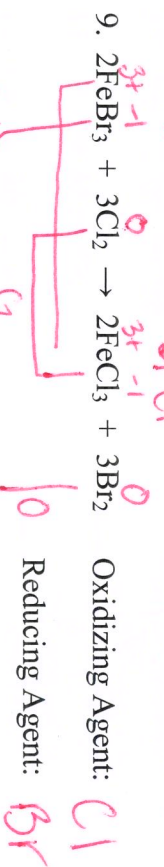
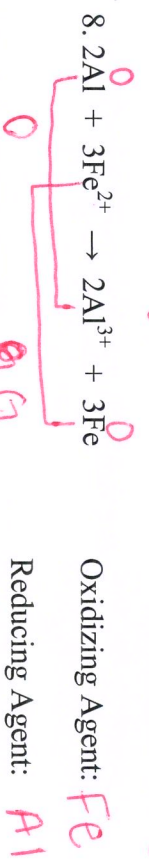
### Redox Mixed Practice WS\_L

**Directions:** Answer the following questions, using your notes and other worksheets, to the best of your ability.

For numbers 1-6, list the oxidation numbers above each element in the following compounds:

- 1.  $\overset{+1}{\text{Li}}\overset{-2}{\text{SO}_4}$
- 2.  $\overset{+5}{\text{N}}\overset{-2}{\text{O}_5}$
- 3.  $\overset{+1}{\text{Na}}\overset{-2}{\text{S}}\overset{-2}{\text{O}_3}$
- 4.  $\overset{+4}{\text{C}}\overset{-2}{\text{O}_3}{}^{-2}$
- 5.  $\overset{+5}{\text{P}}\overset{-1}{\text{Cl}_5}$
- 6.  $\overset{0}{\text{H}_2}$

For numbers 7-10, assign oxidation numbers, identify which element is oxidized and reduced, and list oxidizing agent and reducing agent.



Name: \_\_\_\_\_  
Period: \_\_\_\_\_ Date: \_\_\_\_\_

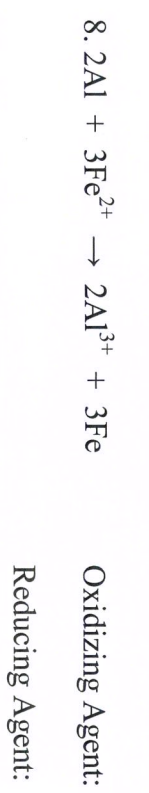
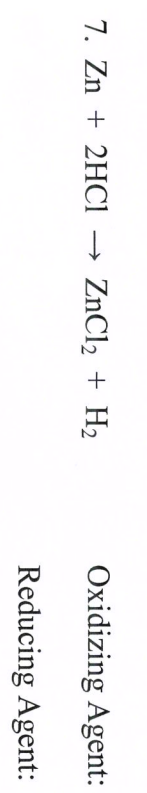
### Redox Mixed Practice WS\_L

**Directions:** Answer the following questions, using your notes and other worksheets, to the best of your ability.

For numbers 1-6, list the oxidation numbers above each element in the following compounds:

- 1.  $\text{Li}_2\text{SO}_4$
- 2.  $\text{N}_2\text{O}_5$
- 3.  $\text{Na}_2\text{S}_2\text{O}_3$
- 4.  $\text{CO}_3{}^{-2}$
- 5.  $\text{PCl}_5$
- 6.  $\text{H}_2$

For numbers 7-10, assign oxidation numbers, identify which element is oxidized and reduced, and list oxidizing agent and reducing agent.



For 11-14 tell whether the half-reaction is oxidation or reduction.



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For numbers 15-18 tell if the compound is soluble or insoluble.



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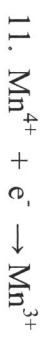


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For number 19 and 20 write the net ionic equation in the box provided.



For 11-14 tell whether the half-reaction is oxidation or reduction.



For numbers 15-18 tell if the compound is soluble or insoluble.



For number 19 and 20 write the net ionic equation in the box provided.

